In summary, we have demonstrated that carbon monoxide can be efficiently trapped by carbon radicals at reasonably low CO pressures, contrary to previous observations. The success of this *free-radical carbonylation* obviously provides a new method for the introduction of carbon monoxide into organic molecules.[9,10] Further studies on the scope and synthetic application are currently under investigtion.

(9) For the preexisting methodologies for carbonylation with CO involving other acyl species (acylmetals, acyl anions, and acyl cations), see recent reviews: (a) Narayama, C.; Periasamy, M. *Synthesis* **1985**, 253. (b) Weil, T. A.; Casser, L.; Foa, M. In *Organic Synthesis via Metal Carbonyls*; Wender, I., Pino, P., Eds.; Wiley: New York, 1977; Vol. 2, p 517. (c) Bahrman, H.; Cornils, B.; Frohling, C. D.; Mullen, A. In *New Syntheses with Carbon Monoxide*; Falbe, J., Ed.; Springer: Berlin, 1980.

(10) The overall transformation as a method for aldehyde synthesis is noteworthy, since the carbonylation of aliphatic substrates is particularly difficult by transition-metal methods. In this context, the radical method reported here complements the method involving Pd catalysis by J. K. Stille; see: Baillargeon, V. P.; Stille, J. K. *J. Am. Chem. Soc.* **1983**, *105*, 7175.

Registry No. 1, 111-83-1; **2**, 124-19-6; **3**, 111-65-9; $Ph(CH_2)_3Br$, 637-59-2; $Ph(CH_2)_2Br$, 103-63-9; $PhCH_2Br$, 100-39-0; PhI, 591-50-4; p-$MeOC_6H_4I$, 696-62-8; c-$C_6H_{11}Br$, 108-85-0; (*E*)-EtCH=CH-$(CH_2)_2Br$, 63281-96-9; (*Z*)-EtCH=CH$(CH_2)_2Br$, 5009-31-4; Ph-$(CH_2)_3CHO$, 18328-11-5; $Ph(CH_2)_2CHO$, 104-53-0; $PhCH_3$, 108-88-3; PhCHO, 100-52-7; p-$MeOC_6H_4CHO$, 123-11-5; c-$C_6H_{11}CHO$, 2043-61-0; (*E*)-EtCH=CH$(CH_2)_2CHO$, 929-22-6; (*Z*)-EtCH=CH-$(CH_2)_2CHO$, 6728-31-0; CO, 630-08-0; 1-adamantyl bromide, 768-90-1; 1-adamantanecarboxaldehyde, 2094-74-8.

Supplementary Material Available: ^1H and ^{13}C NMR, IR, and some mass spectral data for products listed in Table I (3 pages). Ordering information is given on any current masthead page.

Additions and Corrections

Thermal Encapsulation and Photochemical Deencapsulation of Ag(I) by $[Ir_2(dimen)_4](PF_6)_2$ (dimen = 1,8-Diisocyanomethane). X-ray Crystal Structure of $[AgIr_2(dimen)_4](PF_6)_3 \cdot 2DMSO$ [*J. Am. Chem. Soc.* **1988**, *110*, 8252]. ANDREW SYKES and KENT R. MANN*

Page 8253: In Figure 2, the formation constant for the thermal encapsulation of Ag^+ in DMSO is misstated as 1.5×10^8 M^{-1}. The correct value is 1.5×10^7 M^{-1}. The calculations in the text (page 8253) should be adjusted to give log K of 7.2 and the corresponding free energy change of −9.8 kcal/mol.

Time-Resolved Raman Detection of ν(Fe–O) in an Early Intermediate in the Reduction of O_2 by Cytochrome Oxidase [*J. Am. Chem. Soc.* **1989**, *111*, 6439–6440]. CONSTANTINOS VAROTSIS, WILLIAM H. WOODRUFF, and GERALD T. BABCOCK*

The spectrometer used in this characterization of the dioxygen adduct of cytochrome a_3^{2+} in the reaction of fully reduced cytochrome oxidase with O_2 was miscalibrated in the 540–660 cm^{-1} region. With proper calibration, the 589- and 565-cm^{-1} lines occur at 571 and 546 cm^{-1}, respectively. Therefore, we assign the $\nu(Fe^{2+}–O)$ frequency at 571 cm^{-1}. This value is very similar to that which we measured for the iron–oxygen stretching frequency in an imidazole–heme $a^{2+}–O_2$ model compound (see ref 7 in the original publication) and identical with that which we measured in the reaction of mixed valence cytochrome oxidase with O_2. (Varotsis, C.; Woodruff, W. H.; Babcock, G. T. *J. Biological Chem.* Submitted).

This indicates, contrary to our earlier conclusion, that the cytochrome $a_3 \cdot O_2$ complex is unperturbed by distal effects in the cytochrome a_3/Cu_B binding pocket. Weakening and rupture of the O=O bond occurs subsequent to formation of the initial dioxygen–a_3^{2+} adduct.

Selenium Coronands: Synthesis and Conformational Analysis [*J. Am. Chem. Soc.* **1989**, *111*, 6582]. RAYMOND J. BATCHELOR, FREDERICK W. B. EINSTEIN, IAN D. GAY, JIAN-HUA GU, BLAIR D. JOHNSTON, and B. MARIO PINTO*

Recent investigations in our laboratory show that the solid-state ^{77}Se chemical shifts reported in ref 1 are incorrect. This arises from an error in the referencing procedure. Our measurements were referred to a solution of aqueous H_2SeO_3 and converted to the $(CH_3)_2Se$ scale with use of the literature value[2,3] of 1282 ppm for the shift of H_2SeO_3. Direct measurement shows this figure to be incorrect, and the shifts reported in Table XI of ref 1 should be corrected by +22 ppm to give shifts relative to $(CH_3)_2Se$ in $CDCl_3$.

Page 6584: In footnotes *i* and *j* to Table II *W* should be *w*.

Page 6588 first column, line 25: "clockwise" should be "counterclockwise".

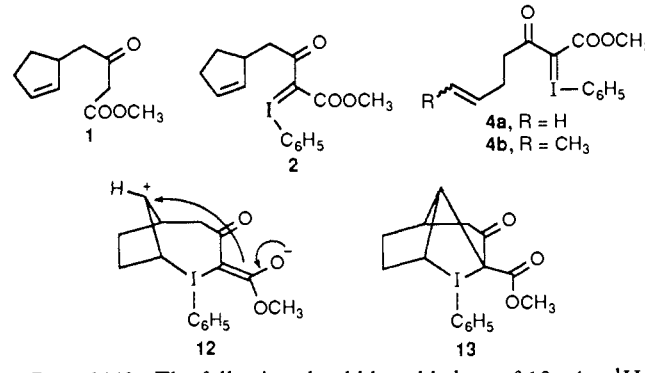
(1) Batchelor, R. J.; Einstein, F. W. B.; Gay, I. D.; Gu, J.-H.; Johnston, B. D.; Pinto, B. M. *J. Am. Chem. Soc.* **1989**, *111*, 6582.

(2) *NMR and the Periodic Table*; Harris, R. K., Mann, B. E., Eds.; Academic Press: New York, 1978.

(3) Lardon, M. *J. Am. Chem. Soc.* **1970**, *92*, 5063.

A Novel Intramolecular Cyclopropanation Using Iodonium Ylides [*J. Am. Chem. Soc.* **1989**, *111*, 6443]. ROBERT M. MORIARTY,* OM PRAKASH, RADHE K. VAID, and LEI ZHAO

Pages 6443 and 6444: The correct structures for **1**, **2**, **4**, **12**, and **13** follow:



Page 6443: The following should be added to ref 10. **4a**: ^1H NMR ($CDCl_3$) δ 2.50 (t, 2 H, CH_2), 3.20 (m, 2 H, CH_2), 3.75 (s, 3 H, $COOCH_3$), 5.10 (m, 2 H, CH=CH_2), 6.09 (m, 1 H, CH=CH_2), 7.40–7.90 (m, 5 H, aromatic protons). **6**: ^1H NMR ($CDCl_3$) δ 1.30–2.10 (m, 4 H, 2 × CH_2), 2.85 (m, 1 H, CH), 3.20 (m, 1 H, CH), 3.70 (s, 3 H, $COOCH_3$), 4.15 (m, 1 H, CH), 5.80 (m, 2 H, CH=CH), 6.20 (m, 1 H, CH=CH), 7.30–7.90 (m, 5 H, aromatic protons).